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1. The saturation vapor
pressure of water at 293 K
is $P_{H_2O, SAT} = 23$ hPa. At
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relative humidity, therefore
 $P_{H_2O} = 11.5 \text{ hPa}$. The dew
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formation 10. 1.2 Phase
partitioning of water in
cloud 10. 1.3 The ozone

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layer 10. 2 ATMOSPHERIC
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atmospheric composition:
mixing ratio, number
density, and partial

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pressure. As we will see,
each measure has its own
applications.

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Introduction to Atmospheric
Chemistry (Princeton
University Press, 1999).

They are arranged following
the different chapters of
the book. In recent years I
have added to my course

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lectures a chapter 14,
'Aerosol Chemistry' and a
chapter 15, 'Mercury in the
Environment'.

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Lecture 1: Introduction I.

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Additional Reading on Units
and Atmospheric Structure

Jacob's chapter 1 on
measures of atmospheric
composition (Suggested)

Jacob's chapter 2 on
atmospheric pressure

(Suggested) Lecture 2:

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Introduction II. Lecture 3:
The Global Atmospheric
Circulation (Brian Toon)

Lecture 4: Atmospheric
Chemical Transport (Brian
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CHEM-5151 / ATOC-5151 -

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Atmospheric Chemistry and
Physics - Exercise A, chap.

1 - 3 Recommended activity
before exercise: Try to
solve 1:1 - 1:5, 2:1 - 2:2
and 3:1 - 3:2. Summary:

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Stability I • Adiabatic
Lapse Rate (?) – vertical
temperature profile when air
ascends or descends
adiabatically, i.e. w/o
giving or receiving heat –
For Earth, $\gamma = 9.8 \text{ K km}^{-1}$ –
Will deduce it from first

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principles later in the
course • Buoyancy force $F_b =$
 $\rho'g - \rho g$ From Jacob

Atmospheric (Vertical)
Stability II

*Lecture 1: Introduction to
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Jacob begins with atmospheric structure, design of simple models, atmospheric transport, and the continuity equation, and continues with geochemical cycles, the greenhouse effect, aerosols,

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stratospheric ozone, the oxidizing power of the atmosphere, smog, and acid rain.

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The objective of atmospheric

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Atmospheric Chemistry is to understand the factors that control the concentrations of chemical species in the atmosphere.

In this book we will use three principal measures of atmospheric composition: mixing

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ratio, number density,
and partial pressure. As we
will see, each measure has
its own applications.

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Tropospheric chemistry 2: CO

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Reactive (oxidized) nitrogen
chemistry (PDF) L 14: Ozone
pollution (PDF) L 15:

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formation in droplets (PDF)
L 17: Atmospheric aerosol 1:
Size, physics (PDF) L 18:
Atmospheric aerosol 2:
Climate effects (PDF) L 19

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appropriate introductory
level text for this class,
and have usually been
borrowing material for the
class from four or five
other books.

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clear review of our basic understanding of the chemistry of Earth's atmosphere. Peter Hobbs is an eminent atmospheric science teacher, researcher, and author of several well-known textbooks.

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measures the frequency of

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collisions of gas molecules with surfaces and therefore determine the exchange rate of molecules between the gas phase and a coexistent condensed phase. Haze.

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